

Name: _____ Date: _____
Chemistry

Worksheet

Molecular Geometry (VSEPR Theory)



DIRECTIONS: Write the orbital notation for the following:

H	___	___	_____	___	_____	___	_____	_____	___
C	___	___	_____	___	_____	___	_____	_____	___
N	___	___	_____	___	_____	___	_____	_____	___
O	___	___	_____	___	_____	___	_____	_____	___
F	___	___	_____	___	_____	___	_____	_____	___
P	___	___	_____	___	_____	___	_____	_____	___
S	___	___	_____	___	_____	___	_____	_____	___
Cl	___	___	_____	___	_____	___	_____	_____	___
B	___	___	_____	___	_____	___	_____	_____	___
Al	___	___	_____	___	_____	___	_____	_____	___

DIRECTIONS: Draw the Lewis dot symbols for the following: H, C, N, O, F, P, S, Cl, B, Al

DIRECTIONS: Determine the number of bonds possible for the following:

_____ H _____ C _____ N _____ O _____ F
_____ P _____ S _____ Cl _____ B _____ Al

DIRECTIONS: Describe the following.

VSEPR Theory -

DIRECTIONS: Draw and label the six basic molecular shapes.

DIRECTIONS: Draw the Lewis structures for the following: CH_4 , C_2H_4 , BeCl_2 , ClF_3 , SF_4 , H_2O . Then determine the electronic shapes, molecular shapes and polarity.

“Remember in the eyes of average people, average is always considered outstanding.”