

Name: _____ Date: _____
Chemistry

Molecular Geometry (VSEPR Theory)



DIRECTIONS: Write the orbital notation for the following:

H	—	—	—	—	—	—	—	—	—
C	—	—	—	—	—	—	—	—	—
N	—	—	—	—	—	—	—	—	—
O	—	—	—	—	—	—	—	—	—
F	—	—	—	—	—	—	—	—	—
P	—	—	—	—	—	—	—	—	—
S	—	—	—	—	—	—	—	—	—
Cl	—	—	—	—	—	—	—	—	—
B	—	—	—	—	—	—	—	—	—
Al	—	—	—	—	—	—	—	—	—

DIRECTIONS: Draw the Lewis dot symbols for the following: H, C, N, O, F, P, S, Cl, B, Al

DIRECTIONS: Determine the number of bonds possible for the following:

_____ H _____ C _____ N _____ O _____ F
_____ P _____ S _____ Cl _____ B _____ Al

DIRECTIONS: Describe the following.

VSEPR Theory -

DIRECTIONS: Draw and label the six basic molecular shapes.

DIRECTIONS: Draw the Lewis structures for the following: CH_4 , AsCl_3 , HCN , ClF_3 , SF_4 , H_2O , then determine the electronic shapes, molecular shapes and polarity.

“Remember in the eyes of average people, average is always considered outstanding.”