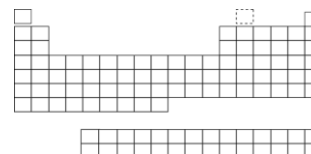


Name: _____ Date: _____
 Chemistry

The Periodic Table II



DIRECTIONS: Read each statement carefully and choose the best term for each statement.

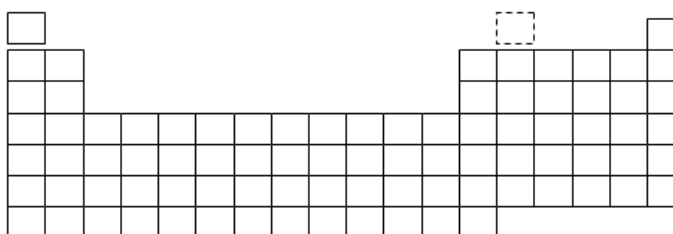
- | | |
|---|------------------------|
| _____ 1. he developed the <u>modern</u> periodic law first | A. atomic number |
| _____ 2. the <u>number of protons</u> in the nucleus | B. families |
| _____ 3. the <u>rows</u> of the periodic table | C. function |
| _____ 4. the electrons in the <u>highest energy level</u> | D. Julius Lothar Meyer |
| _____ 5. the <u>columns</u> of the periodic table | E. Henry Moseley |
| _____ 6. a <u>relation</u> where one set is <u>dependent</u> on the other | F. periods |
| _____ 7. he proposed the <u>first periodic law</u> and table | G. valence |
| | H. Dmitri Mendeleev |

DIRECTIONS: Draw the **Lewis dot symbols** for the following elements.

- | | | | | |
|------------|--------------|--------------|-------------|----------------|
| 8. carbon | 9. strontium | 10. oxygen | 11. boron | 12. chlorine |
| 13. sulfur | 14. copper | 15. nitrogen | 16. calcium | 17. phosphorus |

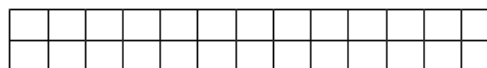
DIRECTIONS: Label the following:

- A. alkaline earth metals
- B. alkali metals
- C. chalcogens
- D. noble gases
- E. halogens
- F. pnictogens



Draw a dark line along the metalloids.

Number the energy levels (1-7)



Label the “s” “p” “d” and “f” blocks

Label the transition metals and the inner transitional metals

***“You don't just stumble into the future. You create your own.”
 -Roger Smith***