

Name: _____ Date: _____
Chemistry

Periodic Trends

DESCRIBE the following terms:

electronegativity –

ionization energy –

atomic radius –

electron affinity –

DIRECTIONS: Tell which ion or atom has the LARGER atomic radius in each of the following sets.

_____ a) F^- , Ne, Na^+ _____ b) O^+ , O, O^- _____ c) Mg^+ , Sr^+ , Ca^+

_____ d) Ca, As, Cr _____ e) F, Kr _____ f) Sr^{2+} , Rb^+ , Kr

DIRECTIONS: Arrange the following elements according to INCREASING first ionization energies.

a) Sr, Be, Mg, Ca, Kr b) F, Te, I, Xe, S

DIRECTIONS: Choose the atom in each of the following groups with the **GREATER** electronegativity.

_____ C, O _____ As, Bi _____ Na, K _____ F, Br _____ I, O

DIRECTIONS: Arrange the following elements according to INCREASING electronegativity.

a) C, N, O, F, Ne b) S, Se, Cl, Ar, Sr

DIRECTIONS: Draw the Lewis dot symbols for the following atoms.

iodine potassium boron cobalt silicon

***"To improve is to change; to be perfect is to change often."
- Winston Churchill***